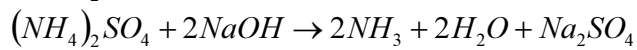
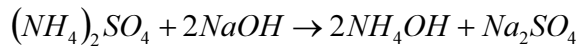


13. Koliko ima molova amonijum – sulfata u jednom litru rastvora, ako 40cm<sup>3</sup> tog rastvora pri reakciji sa natrijum – hidroksidom oslobađa 448cm<sup>3</sup> amonijaka (normalni uslovi)?



$$1mol(NH_4)_2SO_4 \rightarrow 2molNH_3 \quad Vm = 22,4 \frac{dm^3}{mol}$$

$$1mol(NH_4)_2SO_4 \rightarrow 2 \cdot 22,4dm^3NH_3$$

$$xmol(NH_4)_2SO_4 \rightarrow 448 \cdot 10^{-3} dm^3NH_3$$

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$$x = \frac{1mol(NH_4)_2SO_4 \cdot 448 \cdot 10^{-3} dm^3NH_3}{44,8dm^3NH_3}$$

$$x = 0,01mol(NH_4)_2SO_4$$

$$40cm^3rastvora \rightarrow 0,01mol(NH_4)_2SO_4$$

$$1000cm^3rastvora \rightarrow xmol(NH_4)_2SO_4$$

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$$x = 0,25mol(NH_4)_2SO_4$$