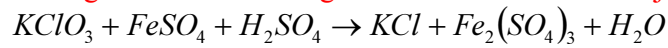
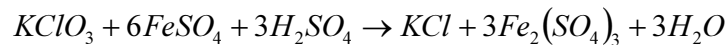
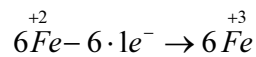
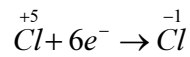
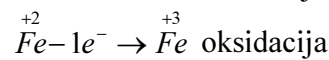
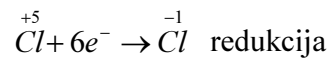
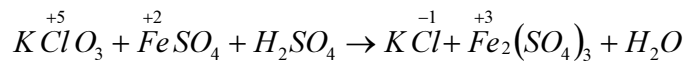


314. Koliko je potrebno grama oksidacionog sredstva da bi se u reakciji:

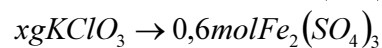
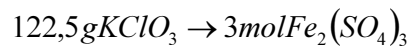
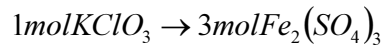


izdvojilo 0,6 mol $Fe_2(SO_4)_3$? ($K = 39; Cl = 35,5; Fe = 56; S = 32$)

Ovo je oksido redukciona reakcija, pri kojoj dolazi do promene oksidacionih brojeva hlora i gvožđa. Hlor se redukovao a gvožđe se oksidovalo.



Kako se hlor redukovao, zaključujemo da je $KClO_3$ oksidaciono sredstvo. Iz reakcije vidimo da:



$$x = 24,5 \text{ g } KClO_3$$